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In electrochemistry, overpotential is the potential difference between a half-reaction's thermodynamically determined reduction potential and the potential at which the redox event is experimentally observed. The term is directly related to a cell's voltage efficiency. In an electrolytic cell the existence of overpotential implies the cell requires more energy than thermodynamically expected to drive a reaction. In a galvanic cell the existence of overpotential means less energy is recovered than

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There have been successes and serendipitous discoveries along the way, but engineering a fuel cell that is both durable and affordable has proved extraordinarily difficult. Overpotential charts the twists and turns in the ongoing quest to create the perfect fuel cell.

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